

**CHM 421/621**

**Statistical Mechanics**

**Lecture 22 Canonical Ensemble**

# Formalisms of Statistical Mechanics

## Lecture Plan

Basic Postulates

Counting microstates in an ideal gas

Boltzmann equation for entropy

Microcanonical Ensemble

Canonical Ensemble

# Ensembles

## Canonical Ensemble

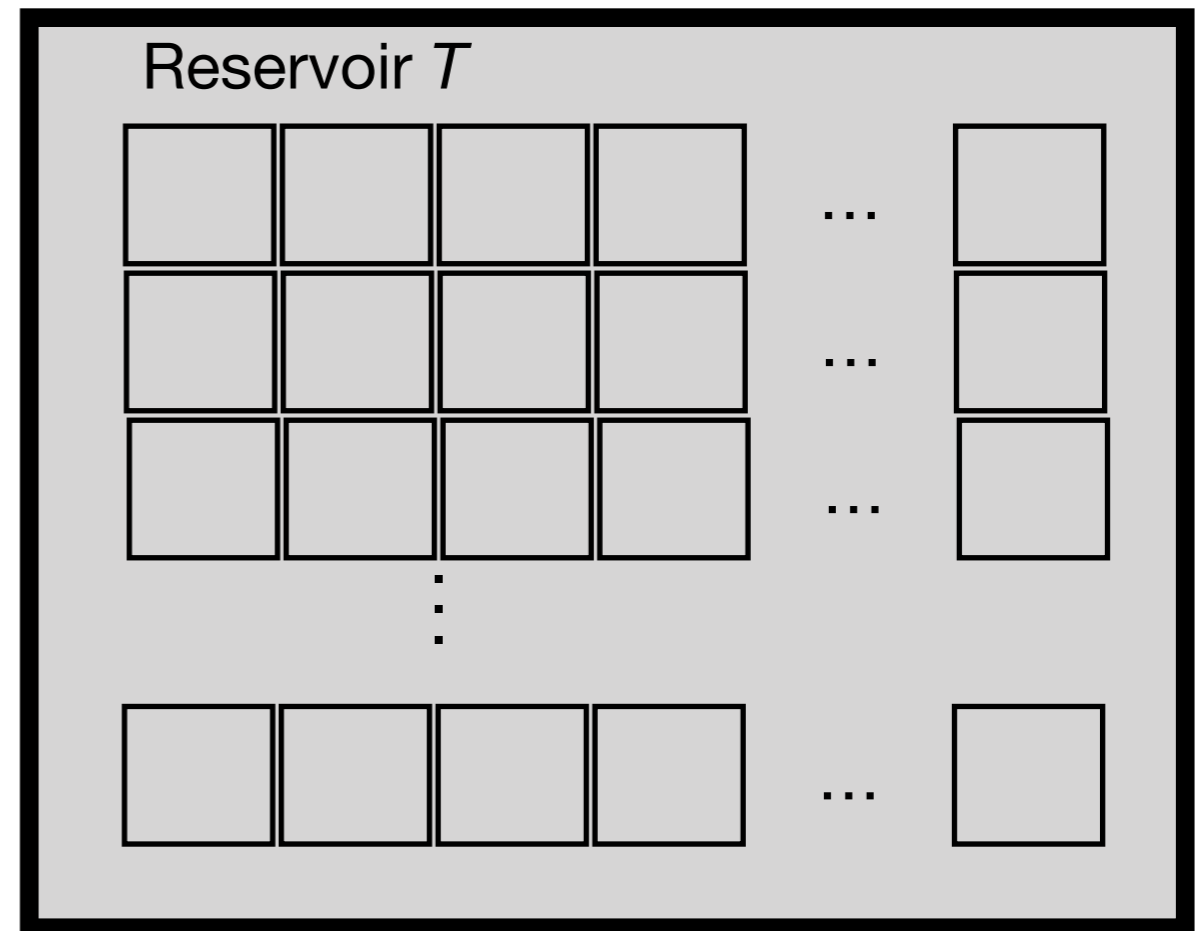
We are generally interested not in isolated systems but those in contact with a thermal (or other) reservoir, capable of exchanging energy with one another.

An ensemble of such (identically prepared) systems at the same temperature ( $T$ ) is called a **canonical ensemble**.

The total energy of the ensemble with reservoir is fixed:  $U_{tot} = \mathcal{E} + U_r$

=> various microstates where  $\mathcal{E}$  and  $U_r$  are different, keeping  $U_{tot}$  constant, are possible.

Even for a given  $\mathcal{E}$  there are multiple ways of distributing the energy across the members of the ensemble. We are actually interested in this distribution.



# Ensembles

## Canonical Ensemble

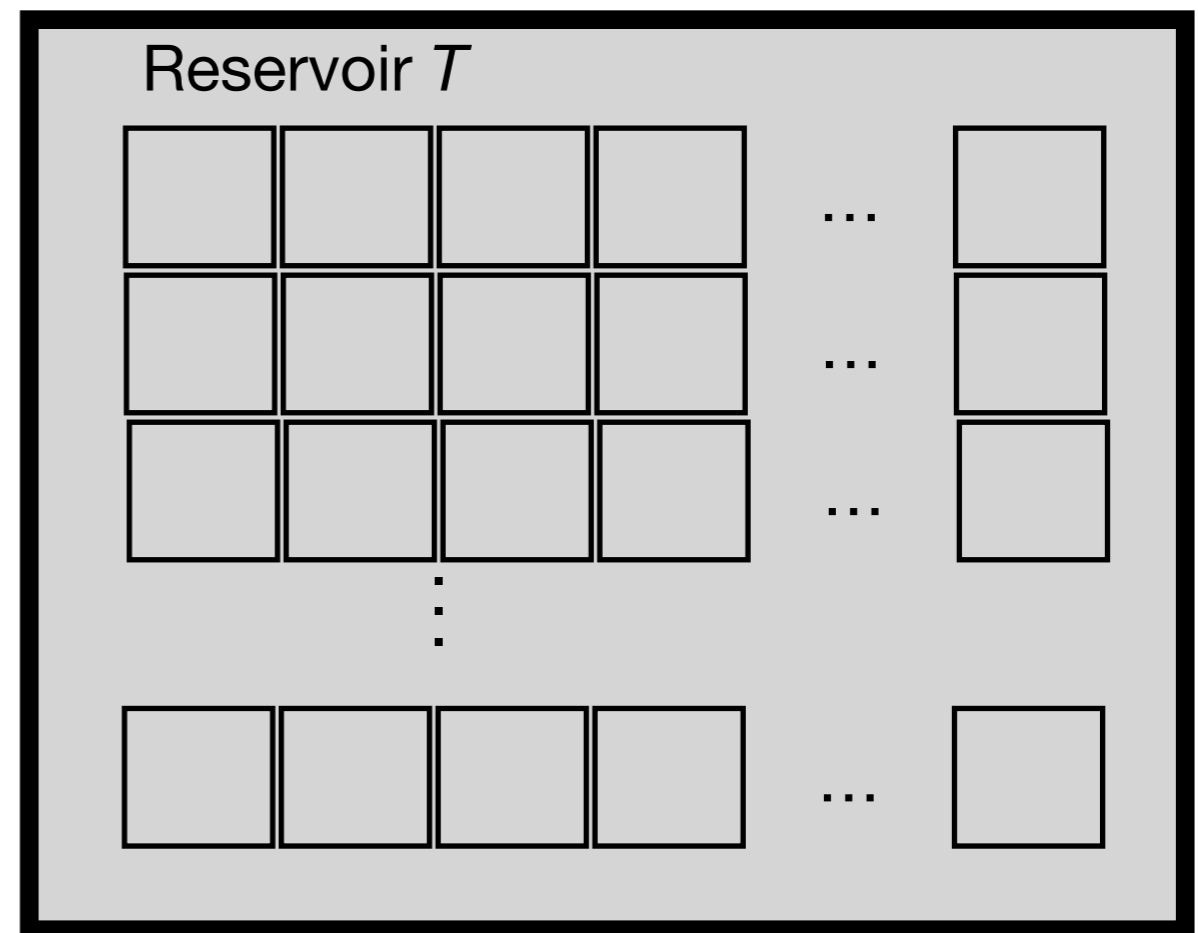
What is the probability that a system in such an ensemble is in a particular microstate  $k$  with energy  $E_k$ ?

From the basic postulate we can see that the probability must depend on the energy of the microstate only. i.e.

Discrete  $P_k \equiv P(E_k)$

Classical  $f(\{\vec{r}, \vec{p}\}) \equiv f(E(\{\vec{r}, \vec{p}\}))$

This can also be understood in terms of grouping the systems into various **micro-canonical ensembles** (hence the name).



# Ensembles

## Canonical Ensemble

What is the probability that a system in such an ensemble is in a particular microstate  $k$  with energy  $E_k$ ?

Joint probability of combined system A+B

$$\begin{aligned} P_{A+B}(E_{A+B}) &\approx P_{A+B}(E_A + E_B) \\ &= P_A(E_A) \times P_B(E_B) \end{aligned}$$

since  $E_A$  and  $E_B$  are independent events.

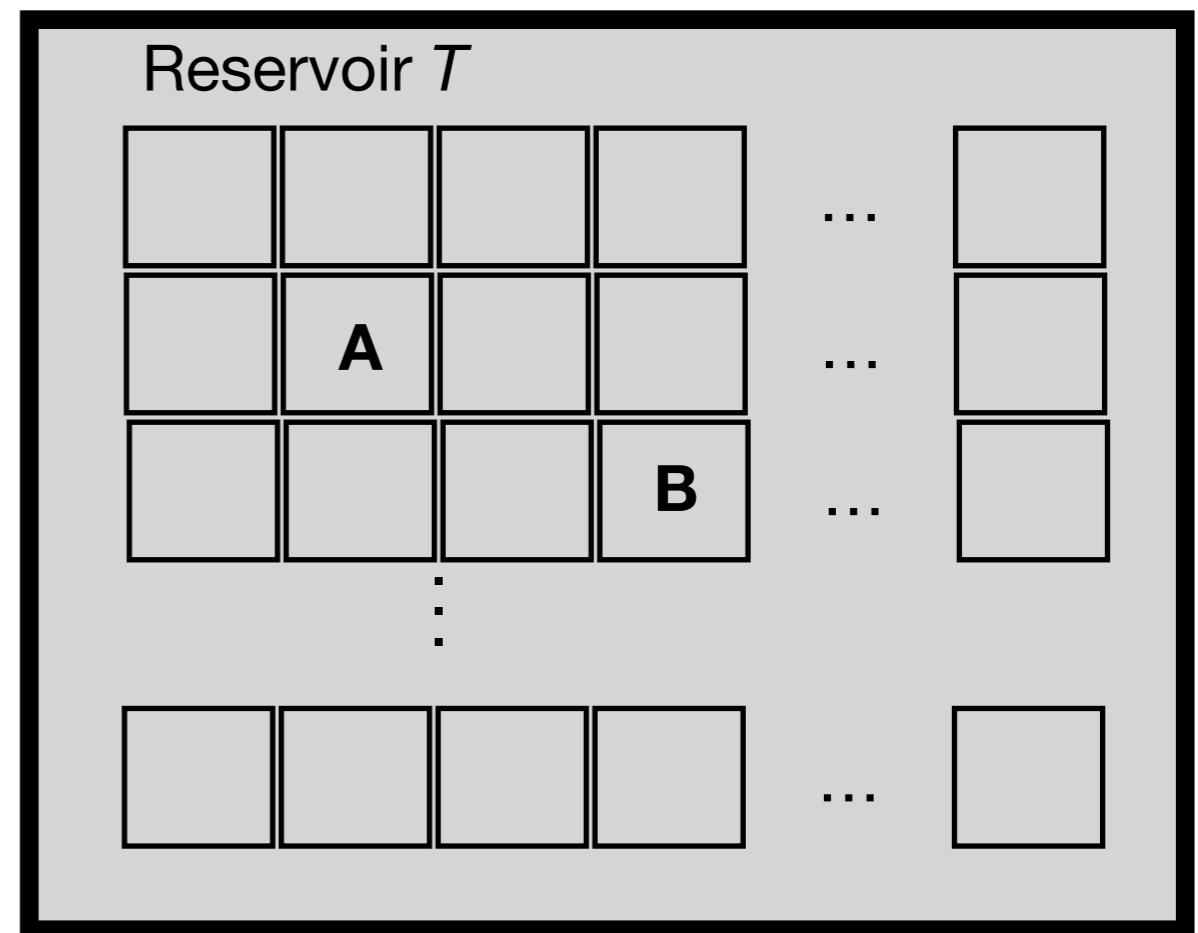
Now,

$$\frac{\partial P_{A+B}(E_{A+B})}{\partial E_A} = \frac{\partial P_{A+B}(E_{A+B})}{\partial E_B}$$

$$\implies P'_A(E_A)P_B(E_B) = P_A(E_A)P'_B(E_B)$$

$$\implies \frac{P'(E)}{P(E)} = \text{constant}$$

Since A and B were arbitrarily chosen.



# Ensembles

## Canonical Ensemble

What is the probability that a system in such an ensemble is in a particular microstate  $k$  with energy  $E_k$ ?

$$\frac{dP(E_k)}{P(E_k)} = C \times dE$$

$C$  is a constant

$$\implies \ln P(E_k) = C E_k + \ln D$$

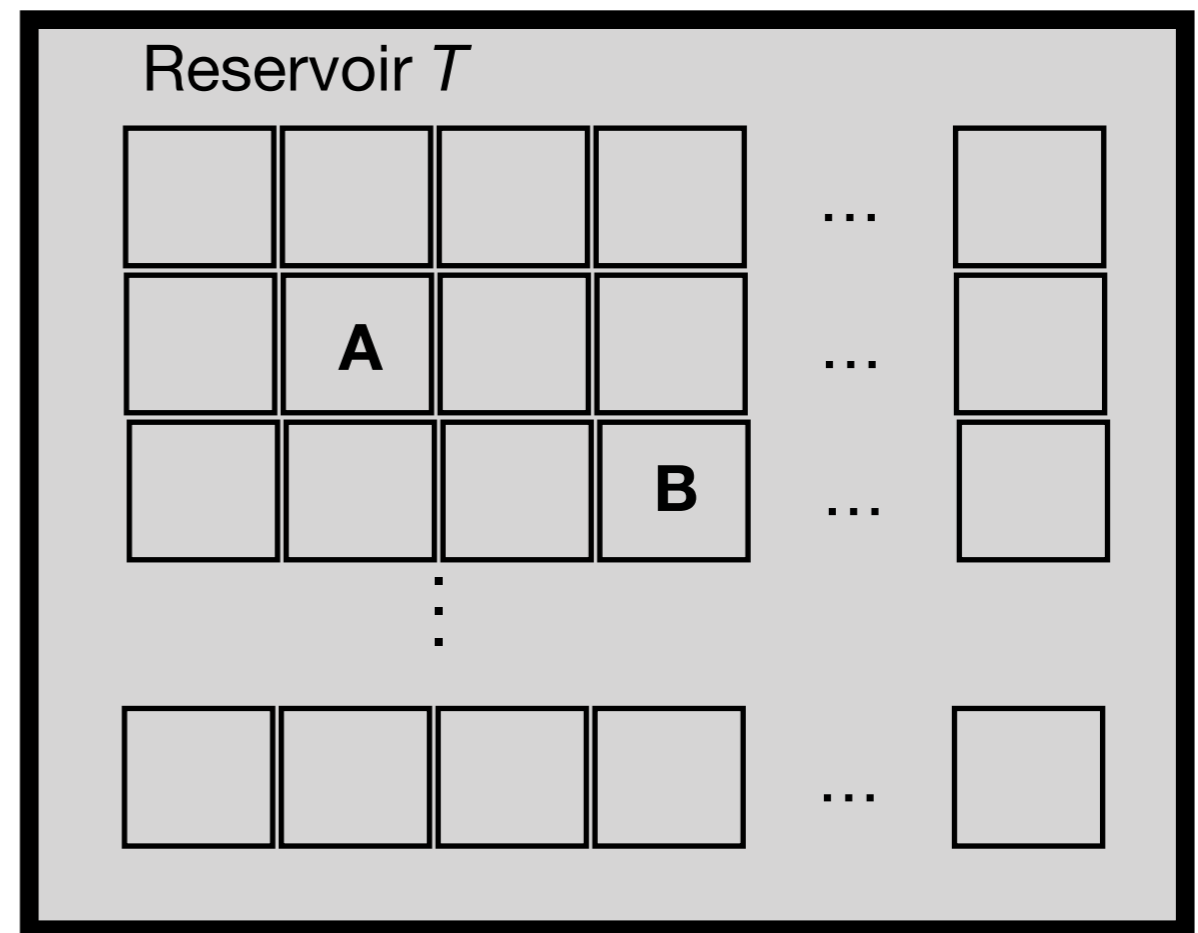
$D$  is a constant of integration

$$\implies P(E_k) = D \exp(C E_k)$$

$$\sum_k P(E_k) = 1$$

$$\implies D = Z^{-1} = 1 / \sum_k \exp(C E_k)$$

$Z$  is called the **canonical partition function**.



# Ensembles

## Canonical Ensemble

What is the probability that a system in such an ensemble is in a particular microstate  $k$  with energy  $E_k$ ?

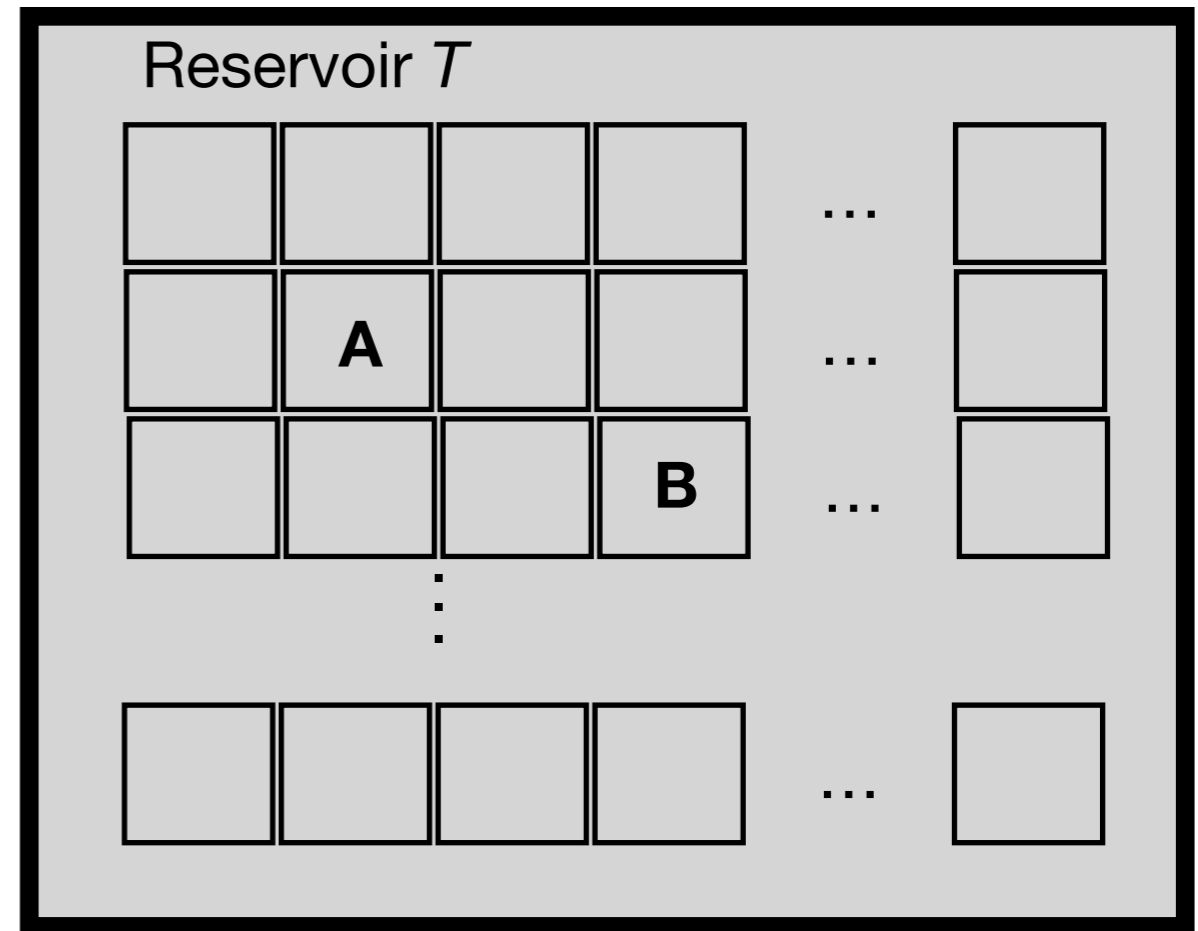
For a discrete (quantum) system

$$P(E_k) = \frac{1}{Z} \exp(-\beta E_k)$$
$$Z = \sum_k \exp(-\beta E_k)$$

(Setting  $C = -\beta$ )

For a continuous (classical) system

$$f(E(\{\vec{r}, \vec{p}\})) = \frac{1}{Z} \exp(-\beta E(\{\vec{r}, \vec{p}\}))$$
$$Z = A \int d^{3N} r \int d^{3N} p \exp(-\beta E(\{\vec{r}, \vec{p}\}))$$
$$A = \frac{1}{N! h^{3N}}$$



$\beta$  must be related to reservoir only!

# Ensembles

## Canonical Ensemble

### Computing averages

$$U \equiv \bar{E} = \frac{1}{Z} \sum_k E_k \exp(-\beta E_k)$$

For classical systems we can write

$$U \equiv \bar{E} = \frac{1}{Z h^{3N}} \int d^{3N}x \int d^{3N}p E \exp(-\beta E)$$

In practice the factor is immaterial as it cancels out (except for the indistinguishability). So we can ignore it by setting it to 1 in the appropriate units.

$$U \equiv \bar{E} = \frac{1}{Z} \int d^{3N}x \int d^{3N}p E \exp(-\beta E)$$



# Ensembles

## Canonical Ensemble

### Computing averages

In either case a simplification to the energy expression is

$$U = -\frac{\partial \ln Z}{\partial \beta}$$

# Ensembles

## Canonical Ensemble

Ideal gas at finite temperature  $T$

$$E = \sum_{i=1}^N p_i^2 / 2m$$

$$\begin{aligned} Z(\beta, V, N) &= \int d^{3N}x \int d^{3N}p \exp(-\beta \sum_{i=1}^N p_i^2 / 2m) \\ &= \left( \int d^3x \int d^3p \exp(-\beta \frac{p^2}{2m}) \right)^N \\ &\equiv (z)^N \end{aligned}$$

where  $z$  is called the molecular partition function.

$$\begin{aligned} z = V \left( \frac{2\pi m}{\beta} \right)^{\frac{3}{2}} &\implies u = U/N = -\frac{\partial \ln z}{\partial \beta} \\ &= \frac{3}{2\beta} \end{aligned}$$

Comparing with the known result of  $u = \frac{3}{2}k_B T$

We have

$$\beta = \frac{1}{k_B T}$$